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A solution is made so that it is 2.5 M $\text{Ca}(\text{NO}_3)_2$. Therefore the concentration of Ca^{2+} is 2.5 M and the concentration of NO_3^- is 5.0 M. The total concentration of all particles is 7.5 M. Explain. Since each mole of $\text{Ca}(\text{NO}_3)_2$ breaks up into 1 mole of Ca^{2+} , the concentration of Ca^{2+} would be the same as the concentration of $\text{Ca}(\text{NO}_3)_2$.

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Chemquest 39 Concentration Answers ChemQuest 39 Name: _____ Date: _____ Hour: _____ Information: Thinking in Quantities A "mole" is simply a quantity of atoms just like a dozen is a quantity of eggs If you have 2 dozen eggs we know you have 24 eggs Similarly, chemists can speak

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Chemquest_39_concentration_answers Chemquest 39 concentration answers| Concentration Practice Worksheet Page 9/30. Get Free Concentration Practice Answers Skills 39 Answers. 1) If I make a solution by adding water to 75 mL of ethanol until the total volume of the solution is 375

Concentration Practice Answers Skills 39

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C1B: ChemQuest 39 Intro to Moles Corrections

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PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys - ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 - HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of Thermodynamics

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Chemquest 1 Numbers In Chemistry Answers

concentration in units of molarity by following these steps: a) Convert the grams of NaCl to moles of NaCl. b) Calculate the litres of solution by dividing the given millilitres by 1000. c) Divide moles by the litres of solution. 4. Calculate the concentration (in units of molarity) of the following solutions: a) 42.1 g of Ca(NO₃)₂

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